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in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of H

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Answer Key Example 1: Calculate the mass of a magnesium, Mg, atoms in grams.  $24.035 \text{ g Mg} \cdot 1 \text{ mol Mg} \cdot 1 \text{ molecule Mg} = 4.04 \times 10^{-23} \text{ g/Mg atom}$   
 $1 \text{ mol Mg} \cdot 6.02 \times 10^{23} \text{ molecules} = 1 \text{ atom Mg}$   
Example 2: Calculate the number of atoms in one-millionth of a gram of

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magnesium, Mg.

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Stoichiometry Review Answers 1. a.  
Na<sub>3</sub>PO<sub>4</sub> b. Ca(NO<sub>3</sub>)<sub>2</sub> Na = 3 mol x 22.99  
g/mol = 68.97 g Ca = 1 mol x 40.08  
g/mol = 40.08 g P = 1 mol x 30.97 g/mol  
= 30.97 g N = 2 mol x 14.01 g/mol =

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28.02 g O = 4 mol x 16.00 g/mol =

64.00 g O = 6 mol x 16.00 g/mol =

96.00 g 163.94 g ...

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equation:  $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$   
How many grams of sodium  
sulfate will be formed if you start with  
200 grams of sodium hydroxide and you

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have an excess of sulfuric acid (H<sub>2</sub>SO<sub>4</sub>)? Using the following equation:  
$$\text{Pb}(\text{SO}_4)_2 + 4 \text{LiNO}_3 \rightarrow \text{Pb}(\text{NO}_3)_4 + 2 \text{Li}_2\text{SO}_4$$

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ANSWER Answer the following questions  
in the space provided. 1. \_\_\_\_\_ The

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